

Write correct chemical names for the following compounds. (1pt each)

- 26. Potassium Chloride KCl I
- 27. Sulfuric Acid H_2SO_4 (aq) #
- 28. Carbon Tetraiodide CI_4 C
- 29. Magnesium Nitrate $Mg(NO_3)_2$ I
- 30. Hydrofluoric acid HF (aq) A
- 31. Aluminum Sulfate $Al_2(SO_4)_3$ I
- 32. Diphosphorus pentoxide P_2O_5 C
- 33. Ammonium Sulfide $(NH_4)_2S$ I
- 34. Chloric acid $HClO_3$ (aq) #
- 35. Silver carbonate Ag_2CO_3
- 36. Calcium hydroxide $Ca(OH)_2$
- 37. dinitrogen tetroxide N_2O_4
- 38. Iron (III) ^{hypo}chlorite $Fe(ClO_3)$
- 39. Barium acetate $Ba(C_2H_3O_2)_2$
- 40. Nitrous acid HNO_2 (aq) O.#
- 41. Tin (IV) Chromate $Sn(CrO_4)_2$ Ionic $Sn^{4+} CrO_4^{2-}$
- 42. Sulfur trioxide SO_3 C
- 43. Copper(I) Phosphate Cu_3PO_4 Cu^+ PO_4^{3-}
- 44. Carbonic Acid H_2CO_3 (aq) MA
- 45. Dinitrogen trisulfide N_2S_3 C
- 46. Manganese(II) Oxide MnO I
- 47. ~~Boron~~ Boron Sulfide B_2S_3 I $B^{+3} S^{2-}$
- 48. Beryllium Peroxide BeO_2 $Be^{2+} O_2^{2-}$
- 49. Hypochlorous acid $HClO$ (aq) ClO^- hypochlorite ^{ous}
- 50. Zinc Oxalate ZnC_2O_4 Ionic $Zn^{2+} C_2O_4^{2-}$

Acids, Covalent Compounds, and Ionic Compounds—Naming and Writing Formulae

Write correct chemical formulas for the following compounds. (1 pt each)

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|-----|-------------------------|------------------------------|
| 1. | silver bromide | AgBr |
| 2. | hydrochloric acid | HCl |
| 3. | dinitrogen trioxide | N_2O_3 |
| 4. | potassium sulfate | K_2SO_4 |
| 5. | iron(III) iodide | FeI_3 |
| 6. | aluminum carbonate | $\text{Al}_2(\text{CO}_3)_3$ |
| 7. | manganese(II) cyanide | $\text{Mn}(\text{CN})_2$ |
| 8. | barium bisulfate | $\text{Ba}(\text{HSO}_4)_2$ |
| 9. | sulfuric acid | H_2SO_4 |
| 10. | calcium nitrite | $\text{Ca}(\text{NO}_2)_2$ |
| 11. | ammonium phosphate | $(\text{NH}_4)_3\text{PO}_4$ |
| 12. | mercury(I) peroxide | Hg_2O_2 |
| 13. | zinc permanganate | $\text{Zn}(\text{MnO}_4)_2$ |
| 14. | phosphorus trichloride | PCl_3 |
| 15. | nitric acid | HNO_3 |
| 16. | tin(IV) hydroxide | $\text{Sn}(\text{OH})_4$ |
| 17. | sodium oxide | Na_2O |
| 18. | silver perchlorate | AgClO_4 |
| 19. | iron(II) sulfide | FeS |
| 20. | mercury(II) bicarbonate | $\text{Hg}(\text{HCO}_3)_2$ |
| 21. | silicon tetrafluoride | SiF_4 |
| 22. | copper(I) carbonate | Cu_2CO_3 |
| 23. | lithium sulfite | Li_2SO_3 |
| 24. | copper(II) nitride | Cu_3N_2 |
| 25. | nickel(II) chlorite | $\text{Ni}(\text{ClO}_2)_2$ |

$\text{Ni}^{2+} \text{ClO}_2^-$
 $\text{Cu}^{2+} \text{N}_3^-$
 $\text{Li}^+ \text{SO}_3^{2-}$
 $\text{Cu}^+ \text{CO}_3^{2-}$
 $\text{Hg}^{2+} \text{HCO}_3^-$

Name _____ Hour _____ Date _____

Key