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| 1. Lead (II) Nitrate solution and zinc metal |
| 1. Iron is mixed with a solution of copper (II) chloride |
| 1. Gold is mixed with a solution of copper (II) chloride |

Please complete the following ASAP.

Determine if each reaction will occur.

If it does, write the complete balanced chemical equation.

Identify which reactant is oxidized and which is reduced.

ANSWERS:

Scoring.

1. States of Matter (SOM)
2. Reaction occurs?
3. Oxidation/reduction
4. Correct equation (reactants, products, balanced)

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| 1. Pb(NO3)2 (aq) + Zn (s) -------> Pb(s) + Zn(NO3)2 (aq)   Pb 2+ = reduced  Zn = oxidized |
| 1. Fe (s) + CuCl2 (aq) -------> Cu(s) + FeCl2(aq)   Cu2+ = reduced  Fe = oxidized |
| 1. CuCl2 (aq) + Au(s) -------> Cu(s) + AuCl (aq)   **No Reaction!** |