

Learning Targets: #7, 10

Name: Key

Hour: _____

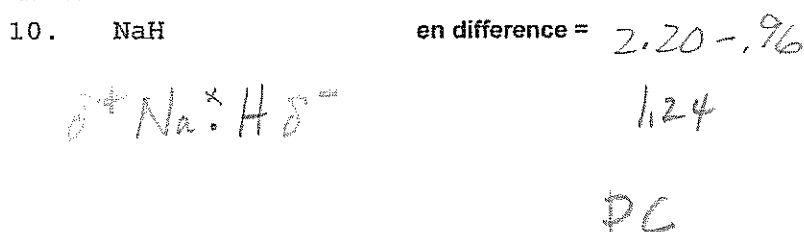
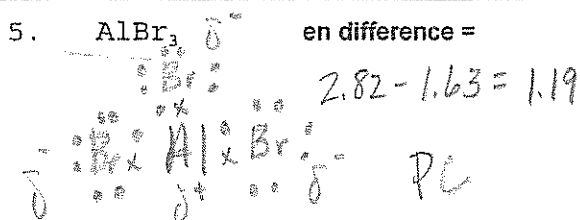
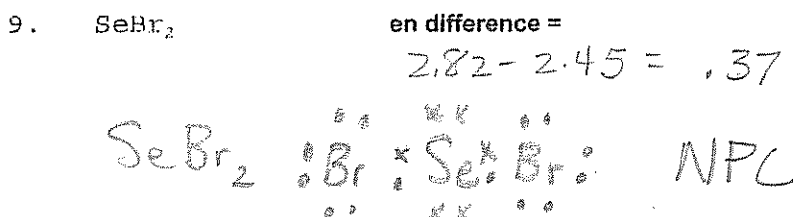
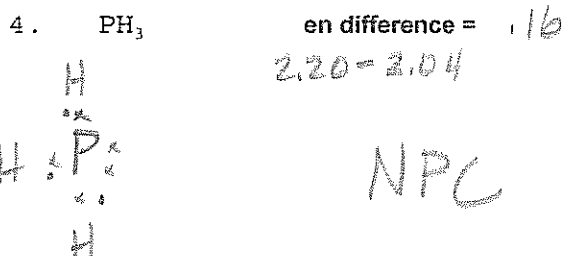
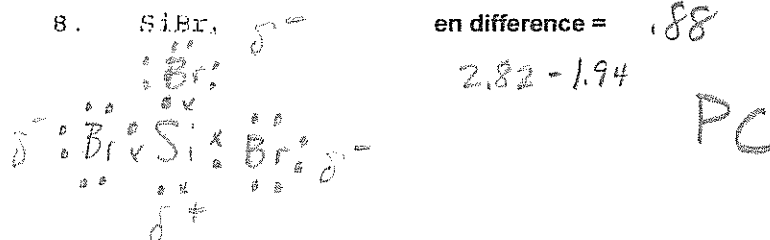
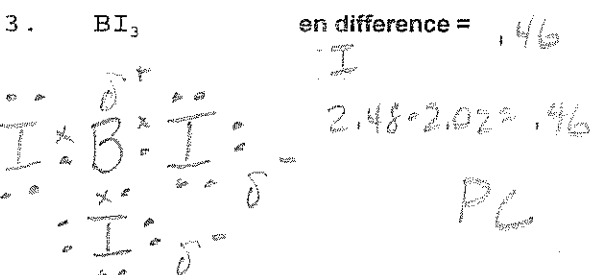
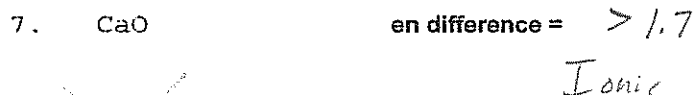
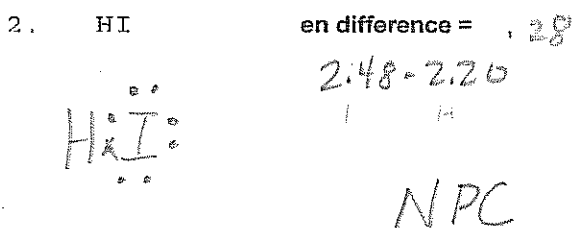
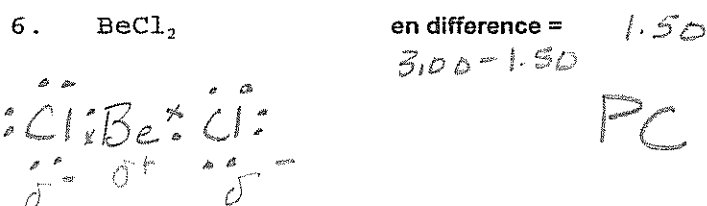
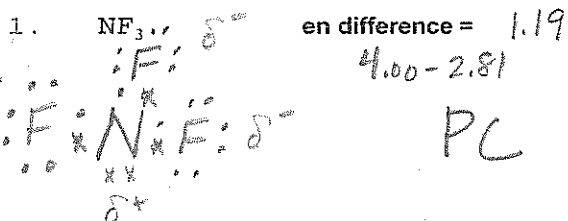
BONDING WORKSHEET #1 - Chemistry

For each of the following compounds:

(A) Use electronegativity values to identify the type of bond as ionic (I), polar covalent (PC), or nonpolar covalent (NPC). If the bonds are polar covalent complete B and C. If they are ionic you only need to write ionic.

(B) DRAW the electron-dot (Lewis) diagram for each of the following compounds.

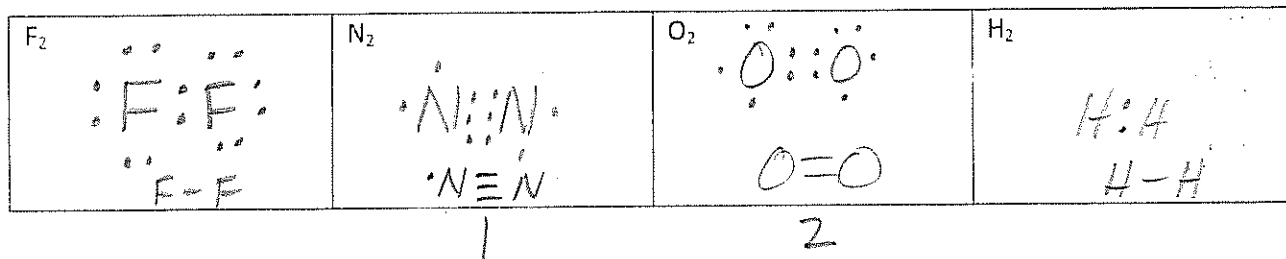
(C) LABEL THE DELTA NEGATIVE & DELTA POSITIVE REGIONS OF THE MOLECULE.



11. Draw electron-dot diagrams for the following molecules below: F_2 , O_2 , N_2 , and H_2

Rank them (1-4) from strongest to weakest bond strength.

Label the molecule with the shortest bond length and the longest bond length. Which of these would take the most energy to break? N_2



12. Name the diatomic molecules. H O F Br I N Cl

13. Are they ionically bonded or covalently bonded? Covalent

14. Which of the compounds on the front page (1-10) will dissolve in water?

1, 6, 3, 8, 5, 10