3.2 ppt Note Guide Part 2

Nucleus:

Protons:

-Location:

-Number of protons =

-In \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ atom number of protons = \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Neutrons:

-Location:

Coulomb’s Law:

-Protons form stable nucleus because strong attractive force\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

-Some \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ help stabilize the \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Atomic Number:

-Same for \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

-No \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_elements have the same atomic number

-Also tells you the number of \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ of an element

Mass Number:

-Calculate the number of \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ by subtracting the

\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Mass number – atomic number =

-Mass number can \_\_\_\_\_\_\_\_\_ among atoms of a single element

Atomic Structures Represented by Symbols

Element Name:

Element Symbol:

Atomic Number:

Example-

Mass Number:

Example-

Isotopes:

-Identify isotopes by writing the \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ with a \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ after the name of the element